

Date Planned : __ / __ / __	Daily Tutorial Sheet - 14	Expected Duration : 90 Min
Actual Date of Attempt : __ / __ / __	Level - 3	Exact Duration : _____

*151. Correct order of electron affinity is/are:

- (A) $S > O$ (B) $Al > B$ (C) $Mg > Na$ (D) $P > N$

*152. Which of the following statement(s) is/are correct?

- (A) van der Waal's radius of iodine is more than its covalent radius
 (B) all isoelectronic ions of corresponding elements belong to the same period of the periodic table
 (C) IE of N-atom is higher than that of O-atom, while IE_2 of O-atom is higher than that of N-atom
 (D) The electron affinity of fluorine is greater than that of chlorine

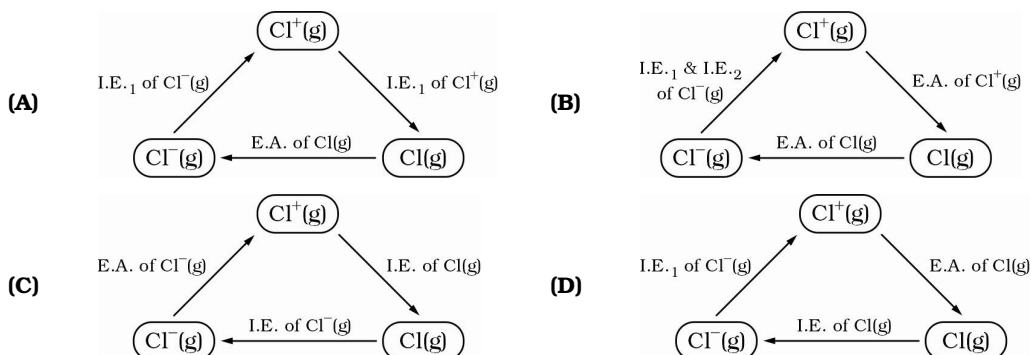
*153. Which of the following is incorrect order of property as indicated?

- (A) $Na^+ < F^- < O^{2-} < N^{3-}$: Ionic size
 (B) $Br < Se < As < Ge$: Metallic character
 (C) $Na < Al < Si < Mg$: Ionisation energy
 (D) $I < Br < Cl < F$: Electron affinity

*154. Which of the following is/are correct order(s)?

- (A) Atomic radius : $F < O < F^- < O^{2-}$
 (B) 2nd ionisation energy : $C < N < F < O$
 (C) Electron affinity : $I < Br < F < Cl$
 (D) Z_{eff} (effective nuclear charge) : $Al < Al^+ < Al^{3+} < Al^{2+}$

155. Select systematic diagram CORRECTLY defines the given change ?



156. Find the total number of cations for which I.P. of cation is lower than corresponding atom.

